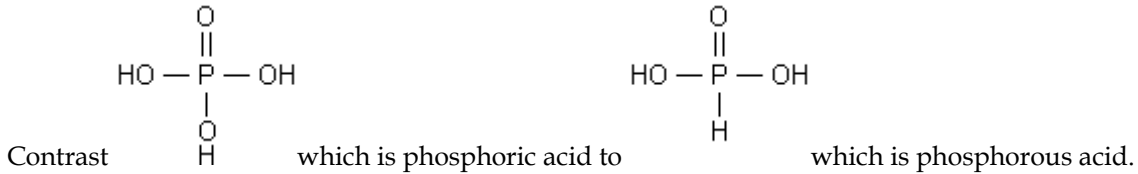


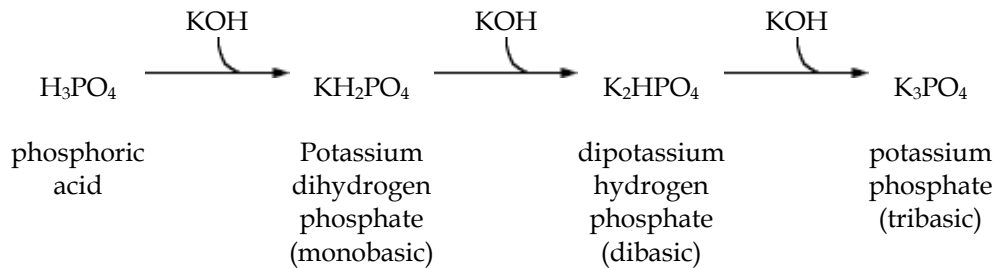
PHOSPHATE VS PHOSPHITE

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CHEMICAL DIFFERENCES OF PHOSPHORIC ACID (Phosphate), PHOSPHOROUS ACID (Phosphite) AND PHOSPHONIC ACID (Phosphonate)

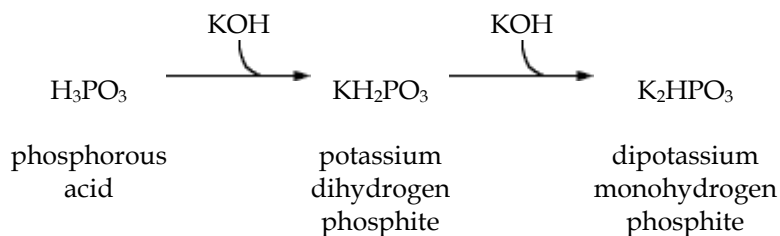


When phosphoric acid (H_3PO_4) is neutralized with a base, such as potassium hydroxide (KOH) or ammonium hydroxide (NH_4OH), a salt results. The salt of phosphoric acid is a phosphate. For example:



(For convenience, all forms of the salt are routinely referred to as "potassium phosphate" with potassium dihydrogen phosphate and dipotassium hydrogen phosphate serving as phosphate fertilizers).

When phosphorous acid (H_3PO_3) is neutralized with a base, such as potassium hydroxide (KOH) or ammonium hydroxide (NH_4OH), a salt results. The salt of phosphorous acid is a phosphite. For example:



An organic derivative of phosphorous acid is a phosphonate. A break down product of a phosphonate is phosphonic acid.